

## CHEMISTRY

### PAPER 3

## CONFIDENTIAL

(a) Each student should be supplied with the following

1. Burette
2. Pipette
3. Pipette filler
4. Filter funnel
5. White tile
6. Clamp and stand
7. 2 conical flask
8. 100 cm<sup>3</sup> glass Beaker (empty)
9. Stop watch
10. 100cm<sup>3</sup> measuring cylinder
11. 10cm<sup>3</sup> measuring cylinder
12. 250 cm<sup>3</sup> volumetric flask
13. Metallic spatula
14. 6 clean test tubes
15. Test tube holder
16. 500ml distilled water
17. White piece of paper or filter paper
18. 1 filter paper
19. 1 labelling paper
20. Phenolphthalein indicator

21. About **90cm<sup>3</sup>** solution **K**
22. About **100cm<sup>3</sup>** solution **L**
23. About **70cm<sup>3</sup>** solution **N**
24. About **90cm<sup>3</sup>** solution **P**
25. About **0.5gNaHCO<sub>3</sub>**
26. About **1.0g** solid **Q**
27. About **0.5g** solid **R**

**(b) Each student should have access to the following solutions:**

1. Mean of heating
2. 2M NaOH
3. 2M HNO<sub>3</sub>
4. Pb(NO<sub>3</sub>)
5. Acidified KMNO<sub>4</sub>
6. Potassium iodide solution

**NB: the above solutions should be supplied with a dropper each.**

**(c) SOLUTIONS PREPARATION AND SOLID MEASUREMENTS**

1. SOLUTION **K** IS 1M H<sub>2</sub>SO<sub>4</sub>
2. SOLUTION **L** IS 0.36 M NaOH CONTAINING 14.4g OF NaOH IN 1 LITRE OF THE SOLUTION
3. SOLUTION **N** IS 2M HCl
4. SOLUTION **P** IS 0.16 M Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>(*SODIUM THIOSULPHATE*)
5. SOLID **Q** IS A MIXTURE OF SODIUM SULPHITE(Na<sub>2</sub>SO<sub>3</sub>) AND LEAD (II) CARBONATE (PbCO<sub>3</sub>) MIXED IN THE **RATIO 1:1** (*should be thoroughly mixed*)
6. SOLID **R** IS MALEIC ACID